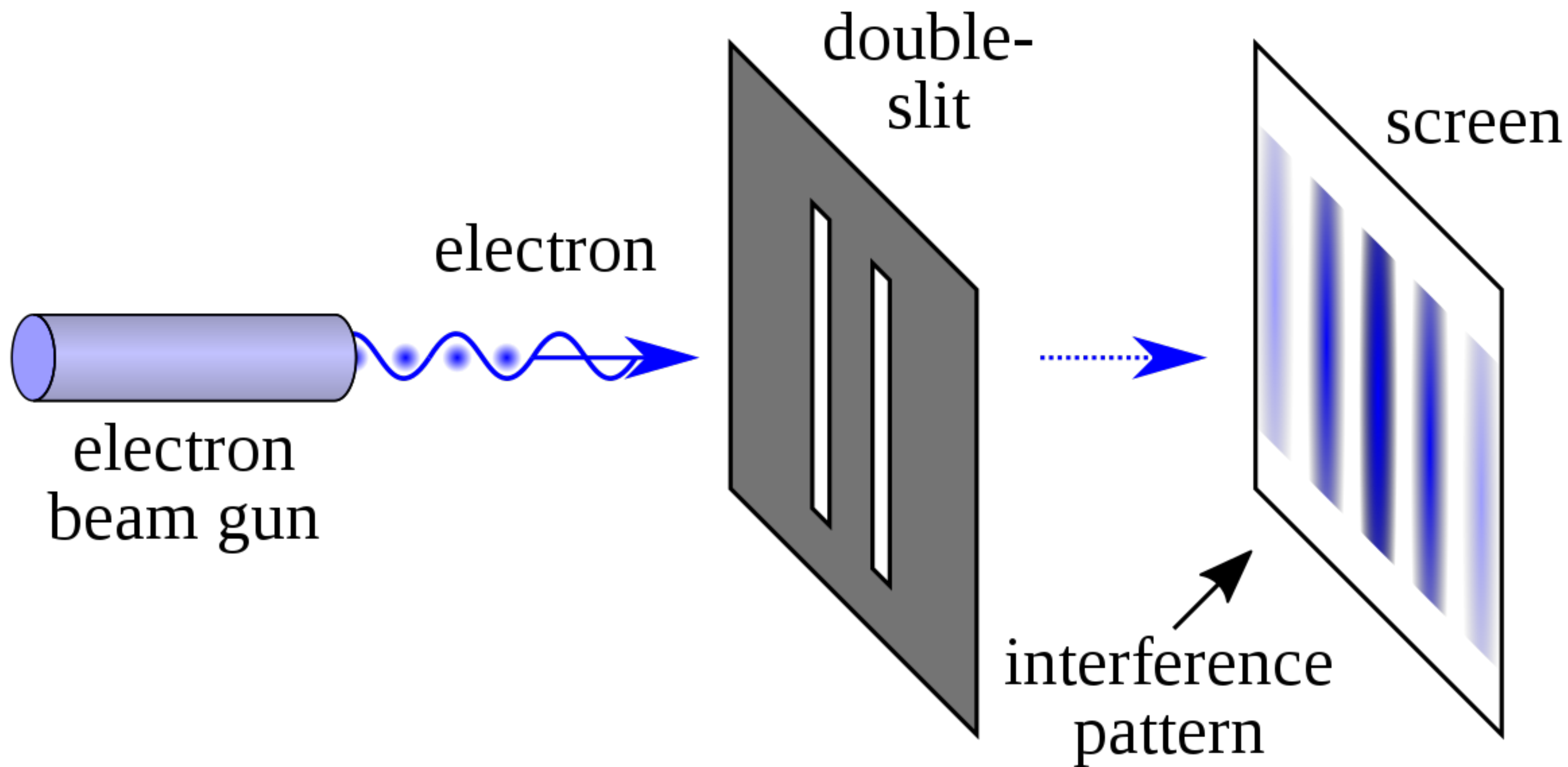


quantum mechanics **(in five minutes)**

brettkoonce.com/talks

october 29th, 2019





Schrodinger equation

$$\hat{H} \Psi = E \Psi$$

Hamiltonian
Operator
(Energy operator)

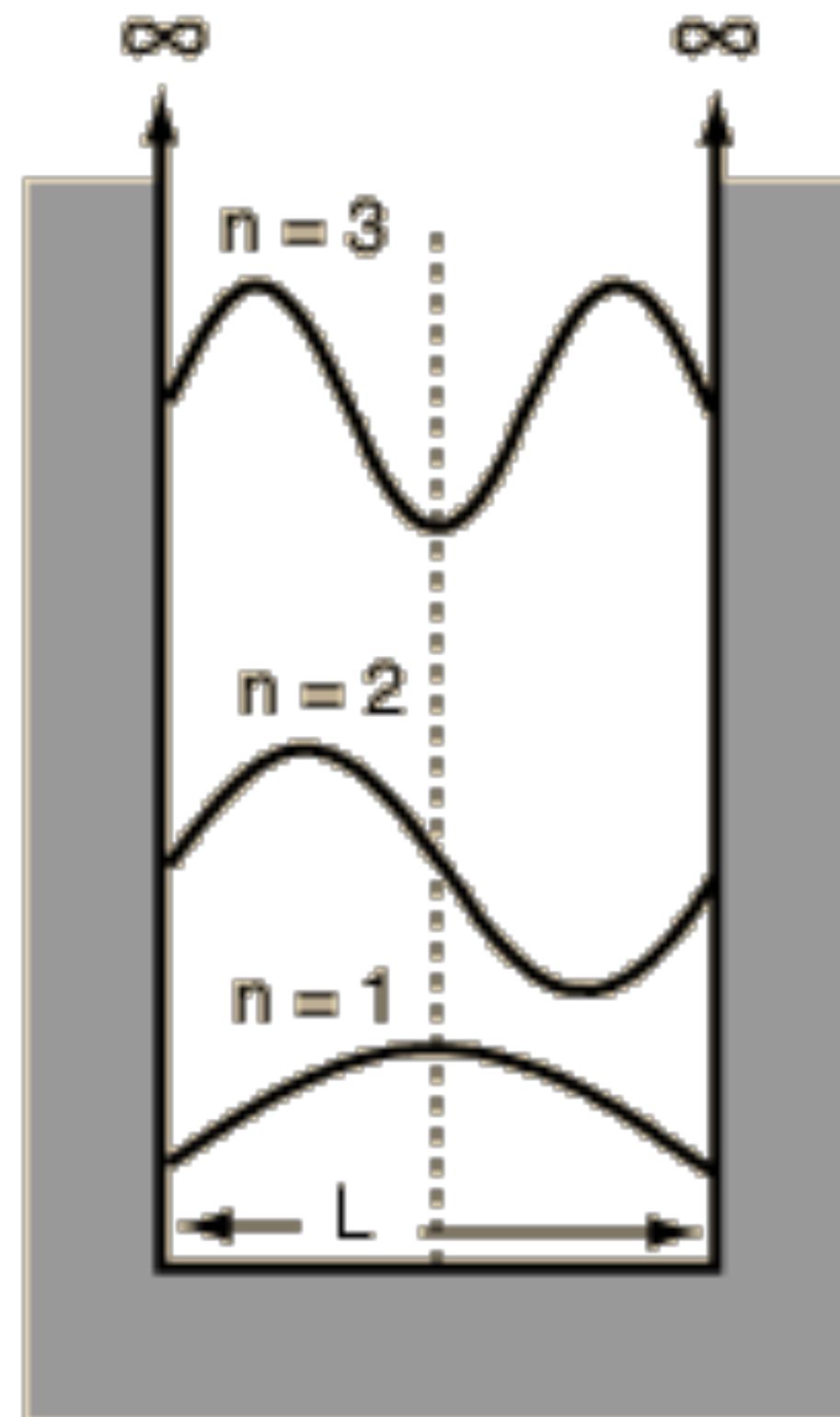
Energy
eigenvalue

Schrodinger equation

$$\frac{-\hbar^2}{2m} \nabla^2 \Psi(\mathbf{r}) + V(r) \Psi(\mathbf{r}) = E \Psi(\mathbf{r})$$

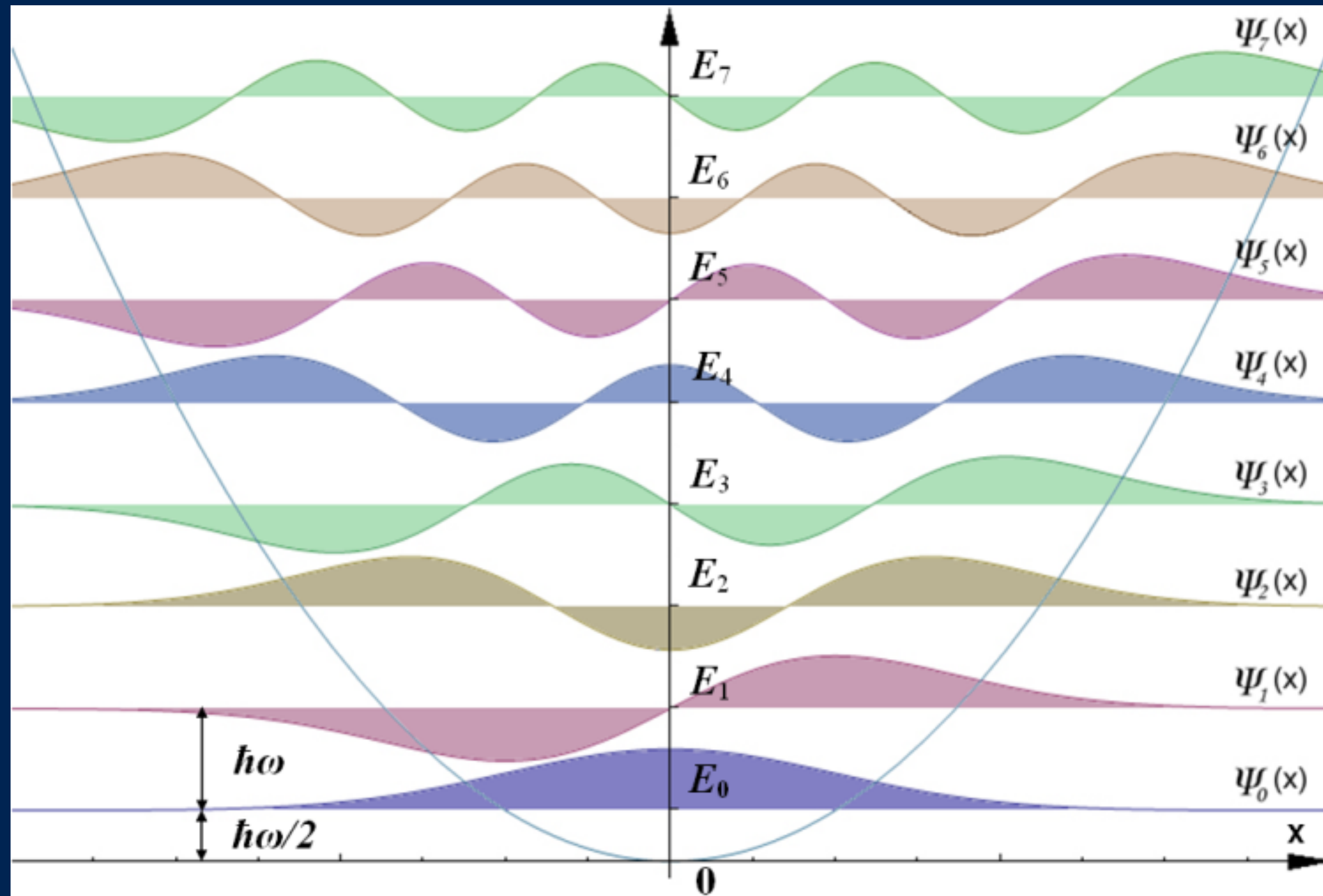
Kinetic Energy + *Potential Energy* = *Total Energy*

particle in a box



$x = 0$ at left wall of box.

1d quantum harmonic



3d form

Second derivative
with respect to Z

Schrodinger wave
Function

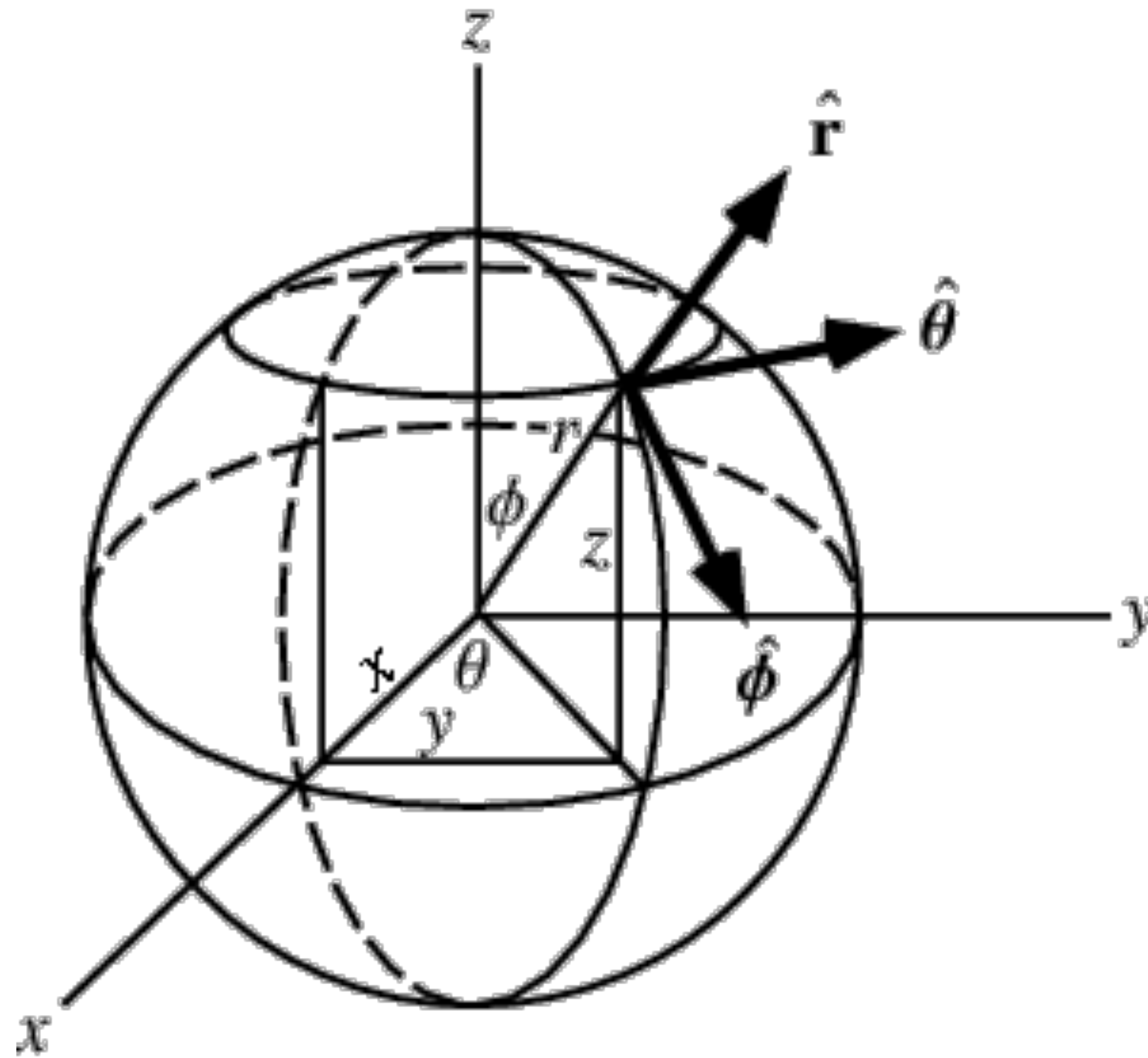
$$\frac{\partial^2 \psi}{\partial x^2} + \frac{\partial^2 \psi}{\partial y^2} + \frac{\partial^2 \psi}{\partial z^2} + \frac{8\pi^2 m}{h^2} (E - V) \psi = 0$$

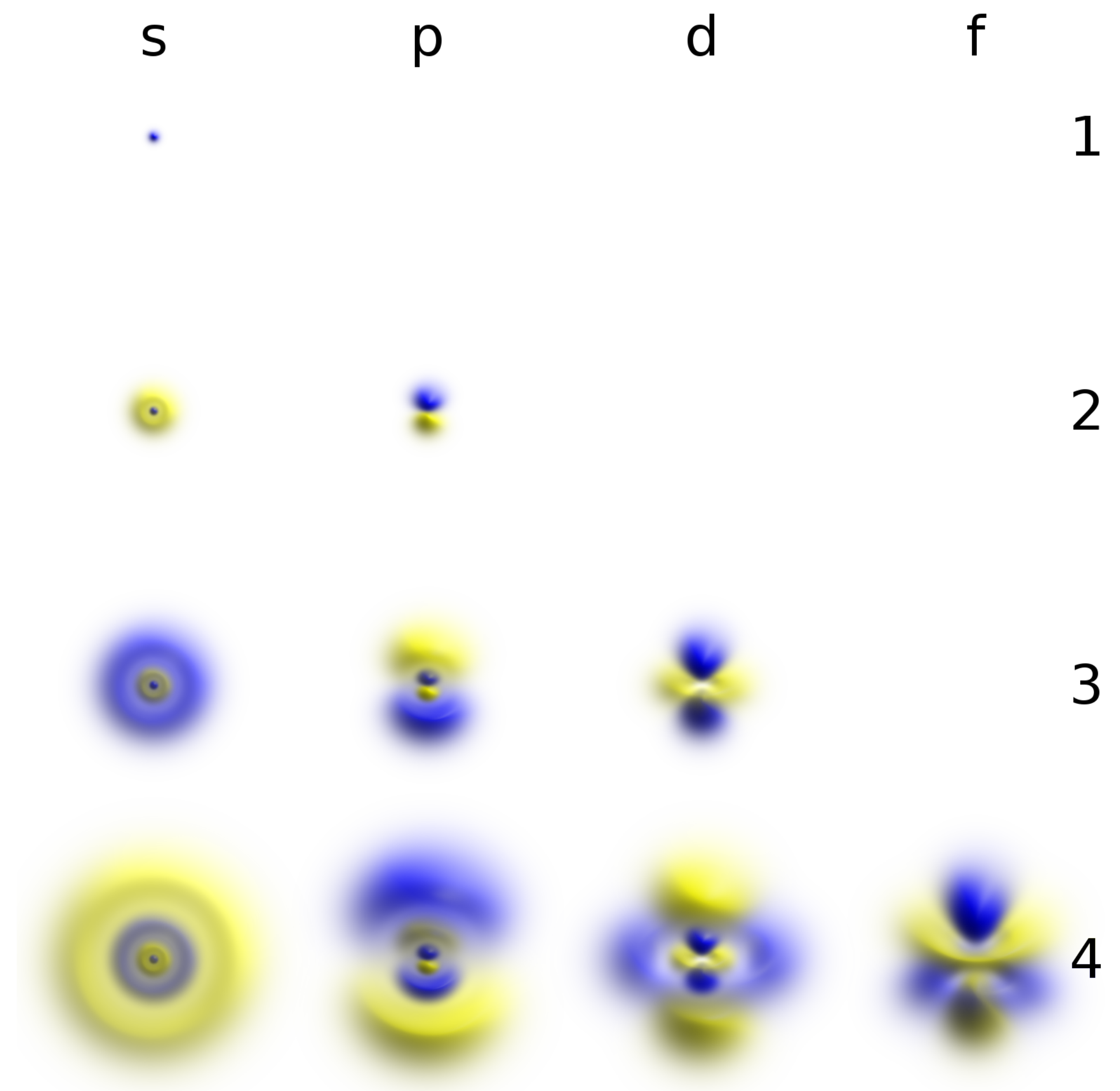
Position

Energy

Potential
Energy

spherical coordinates

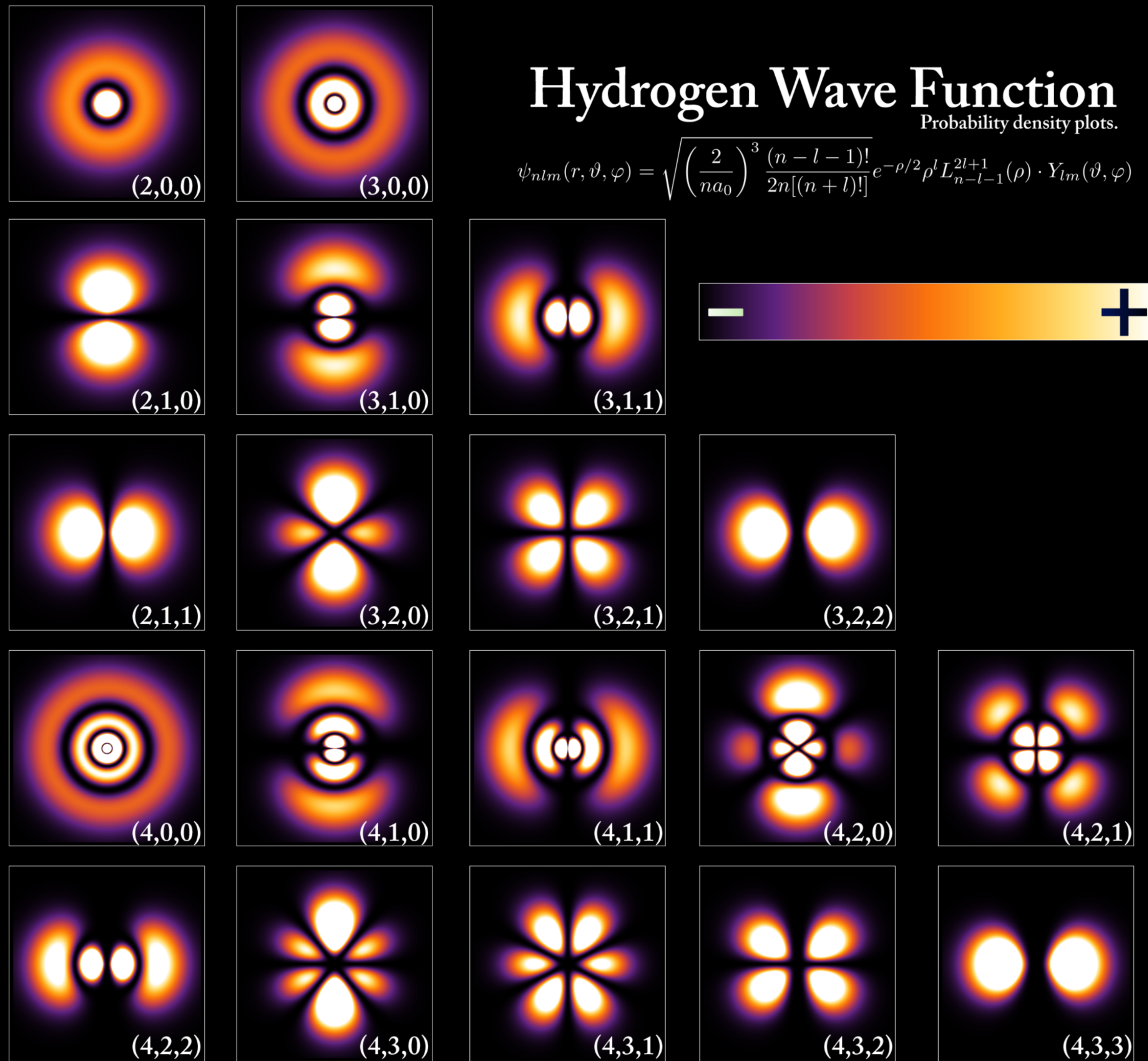




Hydrogen Wave Function

Probability density plots.

$$\psi_{nlm}(r, \vartheta, \varphi) = \sqrt{\left(\frac{2}{na_0}\right)^3 \frac{(n-l-1)!}{2n[(n+l)!]} e^{-\rho/2} \rho^l L_{n-l-1}^{2l+1}(\rho) \cdot Y_{lm}(\vartheta, \varphi)}$$



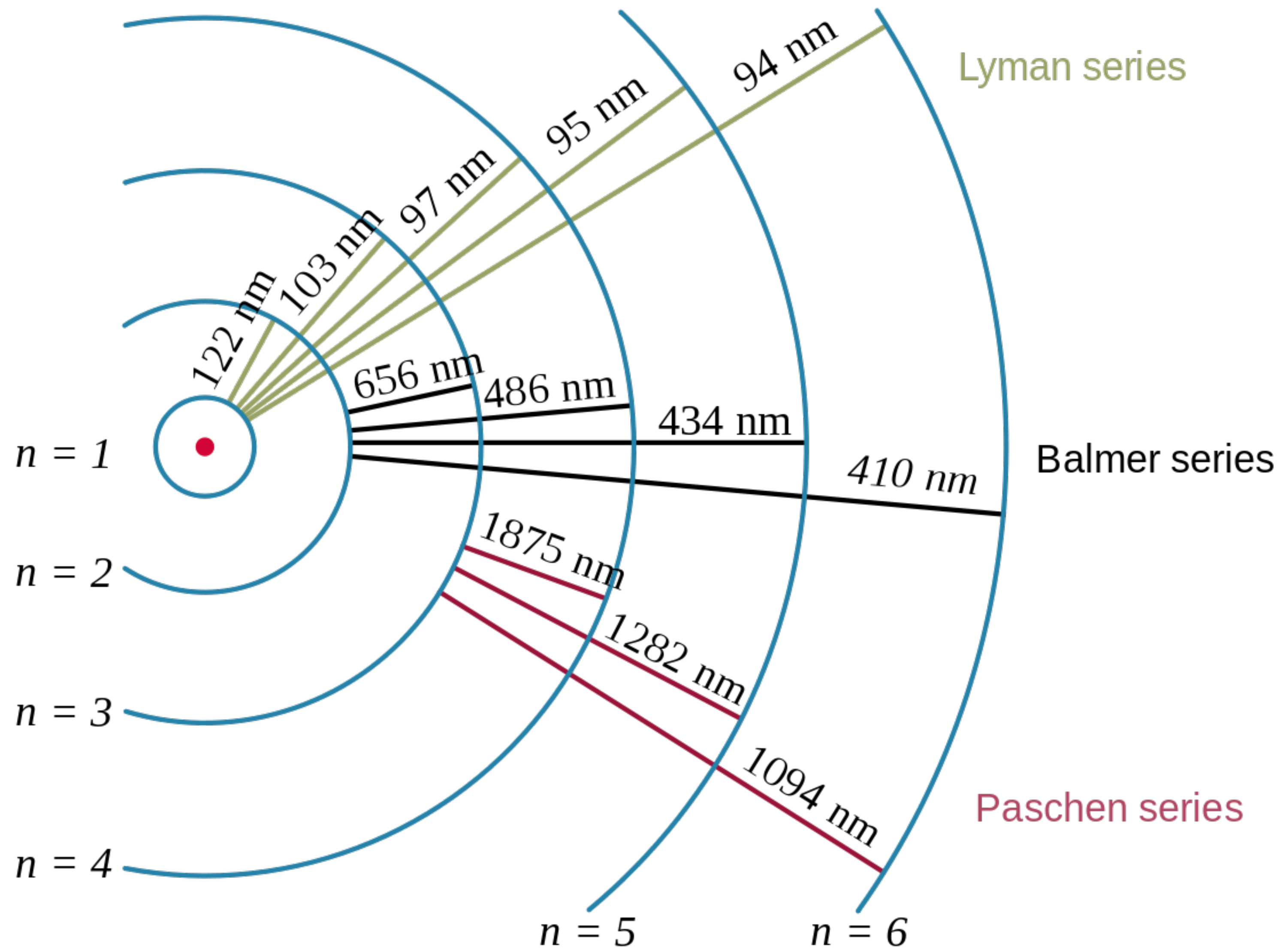
einstein equation

$$E = mc^2$$

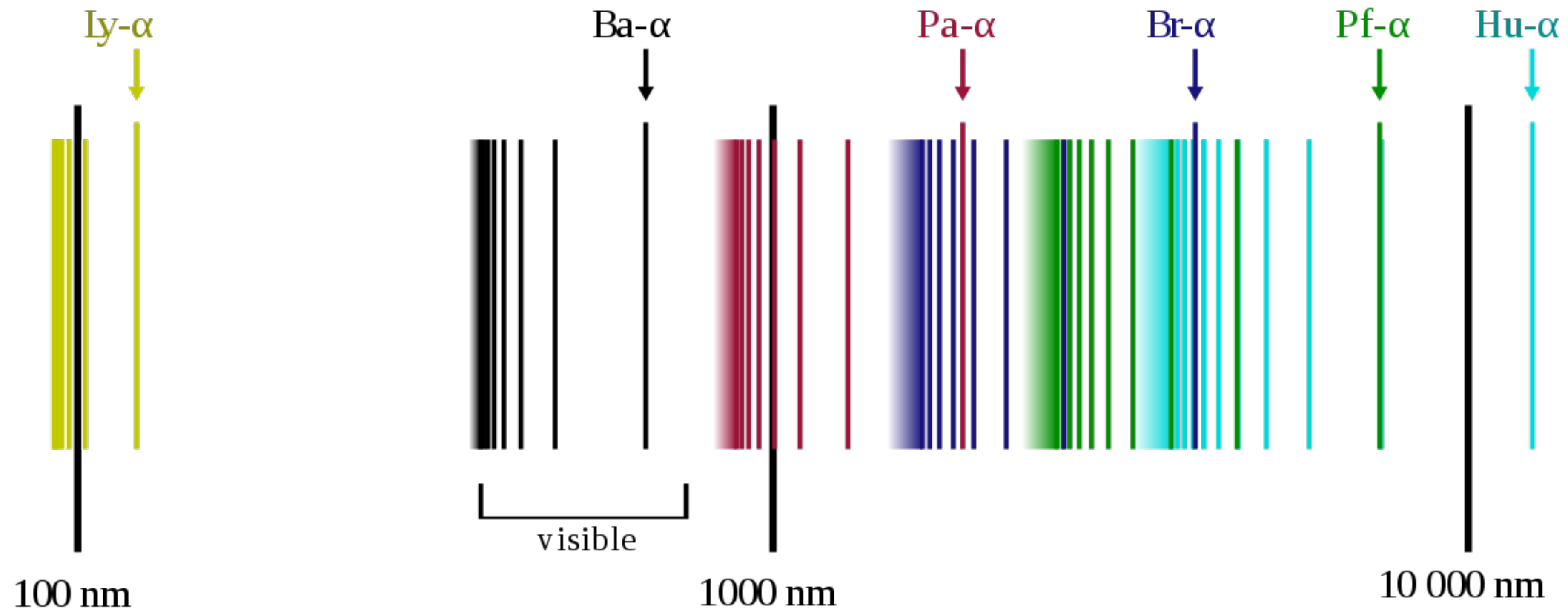


full hydrogen equation

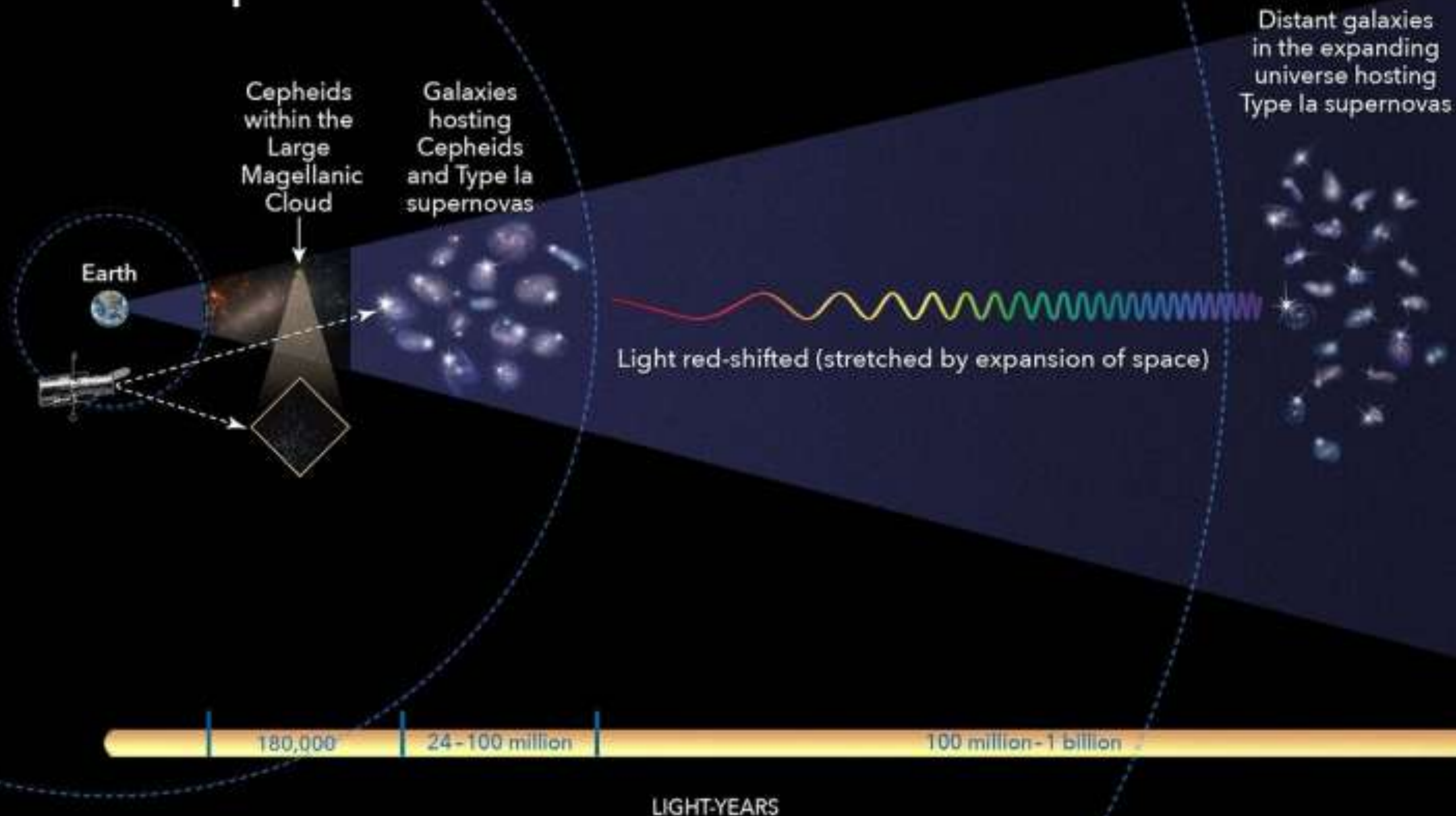
$$\left(\frac{p^2}{2m} - \frac{Ze^2}{4\pi r} - \frac{p^4}{8m^3c^2} + \frac{Ze^2\vec{L} \cdot \vec{S}}{8\pi m^2c^2r^3} + \frac{Ze^2\hbar^2}{8m^2c^2}\delta^3(\vec{r}) \right) \psi = E^{(NR)}\psi$$



hydrogen wavelengths



Three Steps to the Hubble Constant



Bit

(Classical Computing)

0

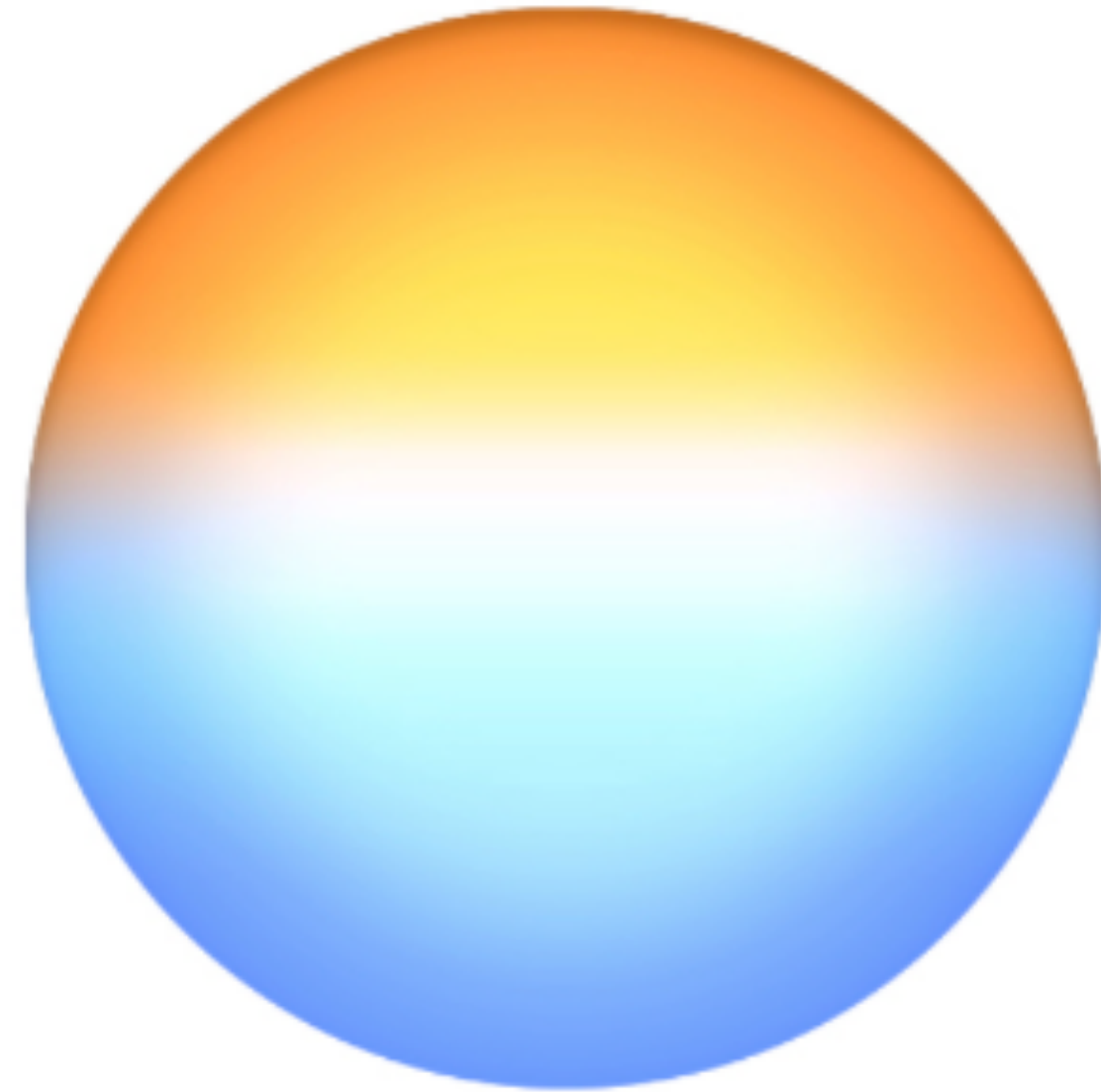


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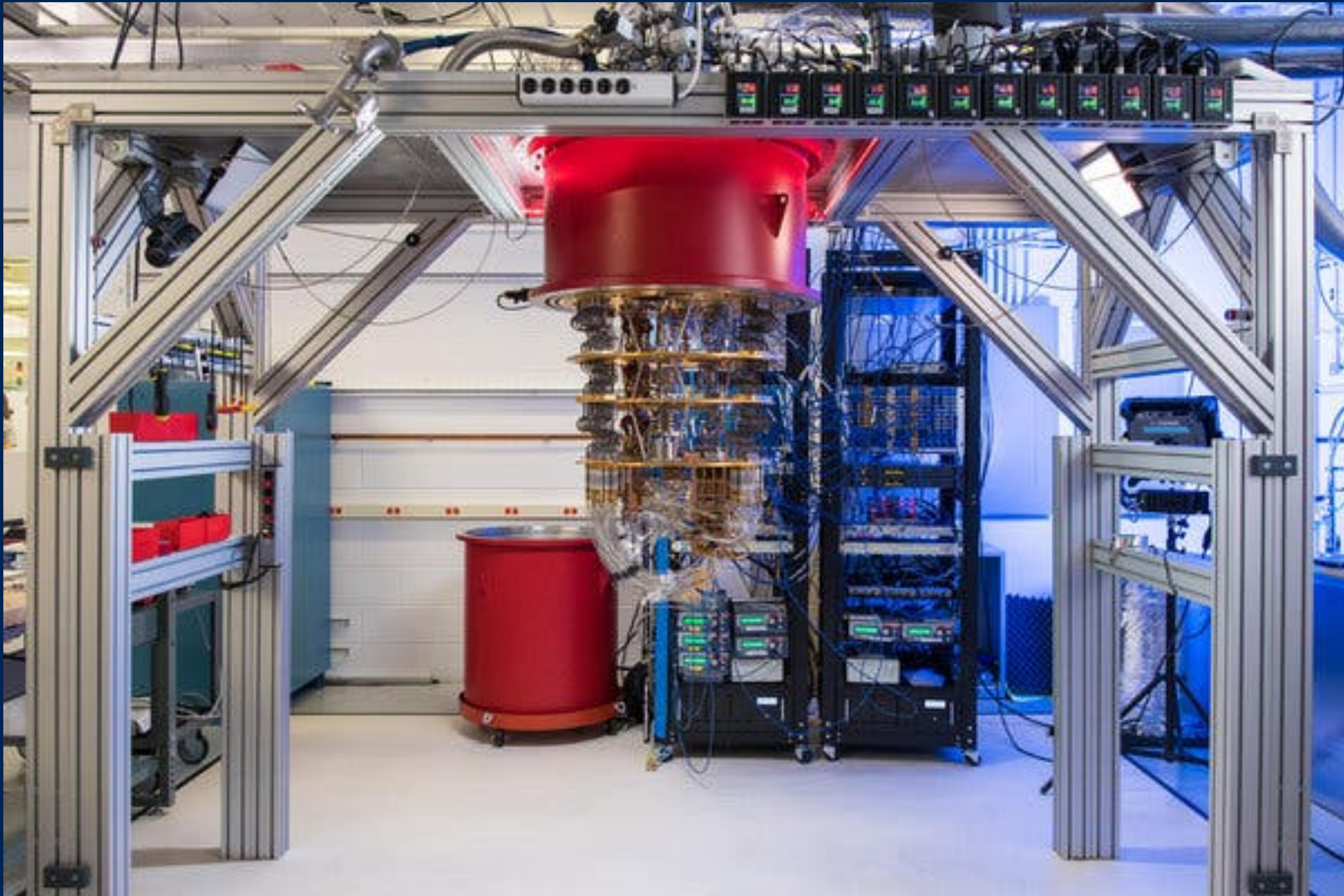
Qubit

(Quantum Computing)

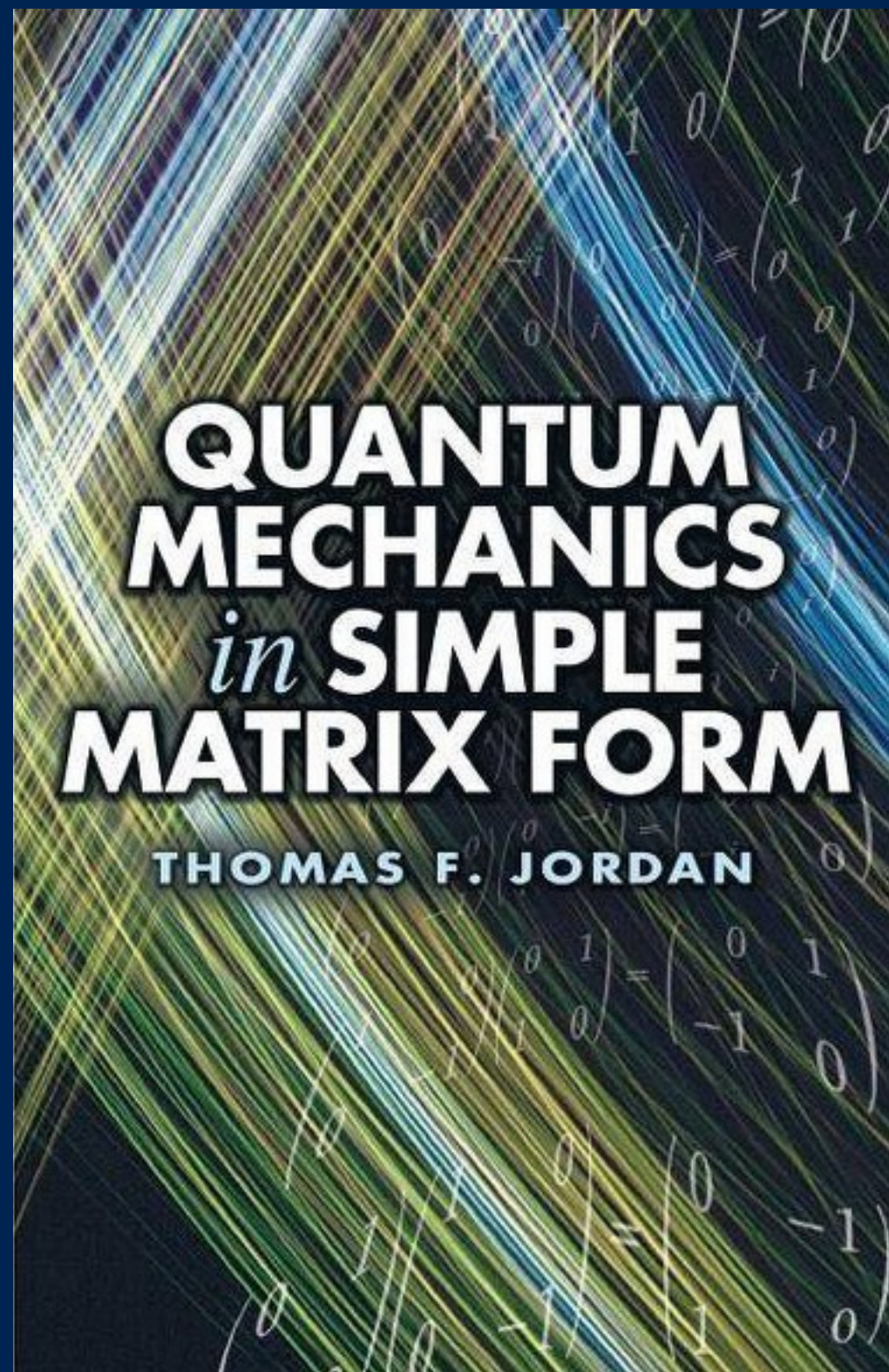
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1



thanks!



- https://quantummechanics.ucsd.edu/ph130a/130_notes/130_notes.html

